The Seven *SI Base Units*

Bureau International des Poids et Mesures

The SI Brochure, Sections 2.2 and 2.3.1

The International System of Units, the SI, is the system of units in which

- **the unperturbed ground state hyperfine transition frequency of the caesium 133 atom, $\Delta\nu_{\mathrm{Cs}}$, is 9 192 631 770 Hz,**
- **the speed of light in vacuum, c, is 299 792 458 m/s,**
- **the Planck constant, h, is $6.626\,070\,15 \times 10^{-34}$ J s,**
- **the elementary charge, e, is $1.602\,176\,634 \times 10^{-19}$ C,**
- **the Boltzmann constant, k, is $1.380\,649 \times 10^{-23}$ J/K,**
- **the Avogadro constant, N_{A}, is $6.022\,140\,76 \times 10^{23}$ mol^{-1},**
- **the luminous efficacy of monochromatic radiation of frequency 540×10^{12} Hz, K_{cd}, is 683 lm/W,**

where the hertz, joule, coulomb, lumen, and watt, with unit symbols Hz, J, C, lm, and W, respectively, are related to the units second, metre, kilogram, ampere, kelvin, mole, and candela, with unit symbols s, m, kg, A, K, mol, and cd, respectively, according to $\mathrm{Hz} = \mathrm{s}^{-1}$, $\mathrm{J} = \mathrm{kg\,m^2\,s^{-2}}$, $\mathrm{C} = \mathrm{A\,s}$, $\mathrm{lm} = \mathrm{cd\,sr}$, and $\mathrm{W} = \mathrm{kg\,m^2\,s^{-3}}$.

The numerical values of the seven defining constants have no uncertainty.

The base units of the SI are listed in Table 2.

Table 2. SI base units

Base quantity		Base unit	
Name	Typical symbol	Name	Symbol
time	t	second	s
length	l, x, r, etc.	metre	m
mass	m	kilogram	kg
electric current	I, i	ampere	A
thermodynamic temperature	T	kelvin	K
amount of substance	n	mole	mol
luminous intensity	I_{v}	candela	cd

Starting from the definition of the SI in terms of fixed numerical values of the defining constants, definitions of each of the seven base units are deduced by using, as appropriate, one or more of these defining constants to give the following set of definitions:

The second

The second, symbol s, is the SI unit of time. It is defined by taking the fixed numerical value of the caesium frequency, $\Delta\nu_{\mathrm{Cs}}$, the unperturbed ground-state hyperfine transition frequency of the caesium 133 atom, to be 9 192 631 770 when expressed in the unit Hz, which is equal to s^{-1}.

This definition implies the exact relation $\Delta\nu_{\mathrm{Cs}} = 9\,192\,631\,770$ Hz. Inverting this relation gives an expression for the unit second in terms of the defining constant $\Delta\nu_{\mathrm{Cs}}$:

$$1\ \mathrm{Hz} = \frac{\Delta\nu_{\mathrm{Cs}}}{9\,192\,631\,770} \quad \text{or} \quad 1\ \mathrm{s} = \frac{9\,192\,631\,770}{\Delta\nu_{\mathrm{Cs}}}.$$

The effect of this definition is that the second is equal to the duration of 9 192 631 770 periods of the radiation corresponding to the transition between the two hyperfine levels of the unperturbed ground state of the $^{133}\mathrm{Cs}$ atom.

The metre

The metre, symbol m, is the SI unit of length. It is defined by taking the fixed numerical value of the speed of light in vacuum, c, to be 299 792 458 when expressed in the unit $\mathrm{m\,s^{-1}}$, where the second is defined in terms of the caesium frequency $\Delta\nu_{\mathrm{Cs}}$.

This definition implies the exact relation $c = 299\,792\,458$ $\mathrm{m\,s^{-1}}$. Inverting this relation gives an exact expression for the metre in terms of the defining constants c and $\Delta\nu_{\mathrm{Cs}}$:

$$1\ \mathrm{m} = \left(\frac{c}{299\,792\,458}\right)\mathrm{s} = \frac{9\,192\,631\,770}{299\,792\,458}\,\frac{c}{\Delta\nu_{\mathrm{Cs}}} \approx 30.663\,319\,\frac{c}{\Delta\nu_{\mathrm{Cs}}}.$$

The effect of this definition is that one metre is the length of the path travelled by light in vacuum during a time interval with duration of 1/299 792 458 of a second.

The kilogram

The kilogram, symbol kg, is the SI unit of mass. It is defined by taking the fixed numerical value of the Planck constant, h, to be $6.626\,070\,15 \times 10^{-34}$ when expressed in the unit J s, which is equal to $\mathrm{kg\,m^2\,s^{-1}}$, where the metre and the second are defined in terms of c and $\Delta\nu_{\mathrm{Cs}}$.

This definition implies the exact relation $h = 6.626\,070\,15 \times 10^{-34}$ $\mathrm{kg\,m^2\,s^{-1}}$. Inverting this relation gives an exact expression for the kilogram in terms of the three defining constants h, $\Delta\nu_{\mathrm{Cs}}$ and c:

$$1\ \mathrm{kg} = \left(\frac{h}{6.626\,070\,15 \times 10^{-34}}\right)\mathrm{m^{-2}\,s}$$

which is equal to

$$1\ \mathrm{kg} = \frac{(299\,792\,458)^2}{(6.626\,070\,15 \times 10^{-34})(9\,192\,631\,770)}\,\frac{h\,\Delta\nu_{\mathrm{Cs}}}{c^2} \approx 1.475\,5214 \times 10^{40}\,\frac{h\,\Delta\nu_{\mathrm{Cs}}}{c^2}.$$

The effect of this definition is to define the unit $\mathrm{kg\,m^2\,s^{-1}}$ (the unit of both the physical quantities action and angular momentum). Together with the definitions of the second and the metre this leads to a definition of the unit of mass expressed in terms of the Planck constant h.

The ampere

The ampere, symbol A, is the SI unit of electric current. It is defined by taking the fixed numerical value of the elementary charge, e, to be $1.602\,176\,634 \times 10^{-19}$ when expressed in the unit C, which is equal to A s, where the second is defined in terms of $\Delta\nu_{\mathrm{Cs}}$.

This definition implies the exact relation $e = 1.602\,176\,634 \times 10^{-19}$ A s. Inverting this relation gives an exact expression for the unit ampere in terms of the defining constants e and $\Delta\nu_{\mathrm{Cs}}$:

$$1\ \mathrm{A} = \left(\frac{e}{1.602\,176\,634 \times 10^{-19}}\right)\mathrm{s^{-1}}$$

which is equal to

$$1\ \mathrm{A} = \frac{1}{(9\,192\,631\,770)(1.602\,176\,634 \times 10^{-19})}\,\Delta\nu_{\mathrm{Cs}}\,e \approx 6.789\,6868 \times 10^{8}\,\Delta\nu_{\mathrm{Cs}}\,e\,.$$

The effect of this definition is that one ampere is the electric current corresponding to the flow of $1/(1.602\ 176\ 634 \times 10^{-19})$ elementary charges per second.

The kelvin

The kelvin, symbol K, is the SI unit of thermodynamic temperature. It is defined by taking the fixed numerical value of the Boltzmann constant, k, to be $1.380\,649 \times 10^{-23}$ when expressed in the unit $\mathrm{J\,K^{-1}}$, which is equal to $\mathrm{kg\,m^2\,s^{-2}\,K^{-1}}$, where the kilogram, metre and second are defined in terms of h, c and $\Delta\nu_{\mathrm{Cs}}$.

This definition implies the exact relation $k = 1.380\,649 \times 10^{-23}$ $\mathrm{kg\,m^2\,s^{-2}\,K^{-1}}$. Inverting this relation gives an exact expression for the kelvin in terms of the defining constants k, h and $\Delta\nu_{\mathrm{Cs}}$:

$$1\ \mathrm{K} = \left(\frac{1.380\,649 \times 10^{-23}}{k}\right)\mathrm{kg\,m^2\,s^{-2}}$$

which is equal to

$$1\ \mathrm{K} = \frac{1.380\,649 \times 10^{-23}}{(6.626\,070\,15 \times 10^{-34})\,(9\,192\,631\,770)}\,\frac{\Delta\nu_{\mathrm{Cs}}\,h}{k} \approx 2.266\,6653\,\frac{\Delta\nu_{\mathrm{Cs}}\,h}{k}.$$

The effect of this definition is that one kelvin is equal to the change of thermodynamic temperature that results in a change of thermal energy kT by $1.380\,649 \times 10^{-23}$ J.

The mole

The mole, symbol mol, is the SI unit of amount of substance. One mole contains exactly $6.022\,140\,76 \times 10^{23}$ elementary entities. This number is the fixed numerical value of the Avogadro constant, N_{A}, when expressed in the unit $\mathrm{mol^{-1}}$ and is called the Avogadro number.

The amount of substance, symbol n, of a system is a measure of the number of specified elementary entities. An elementary entity may be an atom, a molecule, an ion, an electron, any other particle or specified group of particles.

This definition implies the exact relation $N_{\mathrm{A}} = 6.022\,140\,76 \times 10^{23}$ $\mathrm{mol^{-1}}$. Inverting this relation gives an exact expression for the mole in terms of the defining constant N_{A}:

$$1\ \mathrm{mol} = \left(\frac{6.022\,140\,76 \times 10^{23}}{N_{\mathrm{A}}}\right).$$

The effect of this definition is that the mole is the amount of substance of a system that contains $6.022\,140\,76 \times 10^{23}$ specified elementary entities.

The candela

The candela, symbol cd, is the SI unit of luminous intensity in a given direction. It is defined by taking the fixed numerical value of the luminous efficacy of monochromatic radiation of frequency 540×10^{12} Hz, K_{cd}, to be 683 when expressed in the unit $\mathrm{lm\,W^{-1}}$, which is equal to $\mathrm{cd\,sr\,W^{-1}}$, or $\mathrm{cd\,sr\,kg^{-1}\,m^{-2}\,s^3}$, where the kilogram, metre and second are defined in terms of h, c and $\Delta\nu_{\mathrm{Cs}}$.

This definition implies the exact relation $K_{\mathrm{cd}} = 683$ $\mathrm{cd\,sr\,kg^{-1}\,m^{-2}\,s^3}$ for monochromatic radiation of frequency $\nu = 540 \times 10^{12}$ Hz. Inverting this relation gives an exact expression for the candela in terms of the defining constants K_{cd}, h and $\Delta\nu_{\mathrm{Cs}}$:

$$1\ \mathrm{cd} = \left(\frac{K_{\mathrm{cd}}}{683}\right)\mathrm{kg\,m^2\,s^{-3}\,sr^{-1}}$$

which is equal to

$$1\ \mathrm{cd} = \frac{1}{(6.626\,070\,15 \times 10^{-34})\,(9\,192\,631\,770)^2\,683}\,(\Delta\nu_{\mathrm{Cs}})^2\,h\,K_{\mathrm{cd}} \approx 2.614\,8305 \times 10^{10}\,(\Delta\nu_{\mathrm{Cs}})^2\,h\,K_{\mathrm{cd}}\,.$$

The effect of this definition is that one candela is the luminous intensity, in a given direction, of a source that emits monochromatic radiation of frequency 540×10^{12} Hz and has a radiant intensity in that direction of $(1/683)$ W $\mathrm{sr^{-1}}$. The definition of the steradian is given below Table 4.

Text: BIPM, *The International System of Units (SI)*, 9th edition 2019, V4.01 June 2026, Sections 2.2 and 2.3.1, CC BY 4.0; excerpts, equations reset. Compilation: MIT.

The numerical values of the seven defining constants have no uncertainty.

doi.org/10.59161/AUEZ1291 · creativecommons.org/licenses/by/4.0